

Thermodynamic Study of Inclusion Interactions between Gemini Surfactants and β -Cyclodextrin

Xiaomei Qiu*

College of Chemistry and Chemical Engineering, Linyi University, Linyi 276000, China.
E-mail: qiuxiaomeisui@163.com

Received 20 March 2014, revised 21 November 2014, accepted 25 November 2014.

ABSTRACT

The inclusion complexes of gemini surfactants, $(C_nN)_2Cl_2$ (where $n = 12, 14$), and β -cyclodextrin (β -CD) in aqueous solutions were investigated by an isothermal titration calorimetric method and 1H NMR spectroscopy at 293.15 K. The stability constants, stoichiometry, formation enthalpies, entropies and Gibbs energies for the complexes in aqueous solution have been derived from the calorimetric data. β -CD in aqueous $(C_{12}N)_2Cl_2/(C_{14}N)_2Cl_2$ solutions forms 2:1 mole ratio host-guest complexes. The large values of the stability constants indicate that these complexes are stable in the aqueous solutions. The negative Gibbs energy changes indicate that formation of the host-guest complexes is generally a spontaneous process. Both the formation enthalpy and formation entropy evidently decrease as the number of methylenes in each of the hydrophobic tails increase. The thermodynamic parameters are discussed in the light of the different structures of the host and guest molecules. Chemical shift data of protons in the CD molecule, induced by the formation of the complexes have been determined by 1H NMR spectroscopy.

KEYWORDS

Gemini surfactant, β -cyclodextrin, isothermal titration calorimetry, 1H NMR.

1. Introduction

Cyclodextrins (CD) are cyclic oligosaccharides, and their most important properties are their ability to form inclusion complexes with various lipophilic compounds.¹ Surfactant molecules, having a large hydrophobic chain and a polar head group, are one kind of specific guest types for inclusion complex formation.² The complexes of CD and surfactant are widely used as separation and catalysis agents, and in pharmaceutical, food, cosmetic, and agricultural industries.³ Consequently, the inclusion complexes formed between cyclodextrins and surfactants have been studied by means of a number of methods,^{4–7} including conductivity,^{8–10} UV-visible spectroscopy,^{11–12} and surface tension.^{13–19}

Gemini surfactants, which have lately received considerable attention, are molecules that contain more than one hydrophobic tail and hydrophilic head. They usually have better surface activity than that of corresponding conventional ones of equal chain length. They are used as promising surfactants in industrial detergency and have shown efficiency in skin care, antibacterial properties, the construction of high porosity materials, etc.^{20–21}

Previous reports studied the interaction between CD and surfactants and focused mainly on systems containing conventional surfactants.^{5–18} Gemini surfactants were rarely involved. In our work, an isothermal titration calorimeter was used to determine the thermodynamic parameters (stoichiometry, association constant, enthalpy, etc.) for the reaction of gemini surfactants, $(C_nN)_2Cl_2$ ($n = 12, 14$), with β -cyclodextrin at 293.15 K. The results derived from this study have been examined mainly to obtain information about the variation of thermodynamic quantities with different structures of host and guest molecules. Chemical shift data of protons in the CD molecule, determined by 1H NMR spectroscopy, could provide rather detailed information on the formation of complexes.

2. Experimental

2.1. Equipment and Reagents

The experiments were performed in a 201 Nano-Watt Isothermal Titration Calorimeter supplied with a Thermometric 2277 Thermal Activity Monitor (TAM) (Thermometric AB, Sweden). The NMR spectra were obtained with Varian Mercury Plus 400 spectrometer.

The β -CDs were products of Aldrich (purity $\geq 99\%$). The gemini surfactants were synthesized in our laboratory.^{22–25} All solutions were prepared with twice-distilled water for the calorimetric experiments. Deuterium oxide (99.9 at.%) was purchased from Beijing Huagong Chang (Beijing, China).

2.2. Microcalorimetric Measurements

Isothermal Titration Calorimeter has a precision for the electrical calibration of better than $\pm 1\%$. The calorimeter was frequently calibrated electrically during the course of the study and its performance was evaluated by measuring the enthalpy of dilution of a concentrated sucrose solution ($0.985 \text{ mol dm}^{-3}$). The resulting value ($-0.648 \pm 0.012 \text{ kJ mol}^{-1}$) agreed well with the literature value ($-0.653 \text{ kJ mol}^{-1}$).^{26–27} The reference cell of the isothermal titration calorimeter was filled with a fixed amount (2.20 cm^3) of twice-distilled water in order to obtain the heat capacity balance of the system. The sample cell (4.00 cm^3) was filled with 2.00 cm^3 of dilute titrant solution ($(C_nN)_2Cl_2$ solution, $5.00 \times 10^{-4} \text{ mol dm}^{-3}$) and a $500 \mu\text{L}$ Hamilton burette syringe was filled with the titrant solution (β -CD solution, $5.00 \times 10^{-3} \text{ mol dm}^{-3}$) and was controlled by a 612 Lund Syringe Pump. The titration process consisted of 40 aliquots of $12 \mu\text{L}$. The interval between two injections was 0.50 to 0.75 h, which was sufficiently long for the signal to return to the baseline. The golden propeller stirrer in the ampoule was rotated at a constant speed of 50 r.p.m. An injection schedule was automatically carried out that was

controlled by Digitam 4.1 software after programming the number of injections, the volume of each injection and the time between each injection.²⁸ All experiments were performed at 293.15 ± 0.01 K. The dilution enthalpies of the surfactants and the CD solutions were determined for each titration experiment by using the same number of injections and concentrations as in the titration experiments. No significant effect was found so that the dilution heat was negligible.

2.3. NMR Measurements

¹H NMR experiments were carried out in deuterium oxide. The spectra were acquired at 293.15 K. The center of the solvent signal (D₂O, at 4.640 ppm) was used as the reference. The proton chemical shifts of β -CD were determined with the increased concentrations of the surfactants (C_{*n*}N)₂Cl₂.

3. Results and Discussion

3.1. Data Processing

To rationalize the formation enthalpy of the inclusion complexes of β -CD with (C_{*n*}N)₂Cl₂ in aqueous solution, apparent equilibrium constants, β_i , for the overall reactions have been defined as:



where $i = 1$ or 2 or 3, meanwhile M or L can represent either the host (β -CD) or guest ((C_{*n*}N)₂Cl₂). The relationship between the overall equilibrium constants, β_i , and the stepwise equilibrium constants, K_i , is:

$$\beta_i = \prod K_i \quad (2)$$

The thermodynamic parameters, ΔH_i° and K_i° , can be calculated by using a regression method,²⁹ which was performed with the Ligand Binding Analysis process using Digitam 4.1 software.²⁸ We obtained optimum results by comparison of the simulated (calculated) curve with the experimental points.³⁰ The optimum simulated results of the supramolecular systems are shown in Figs. 1 and 2. Thermodynamic parameters for the inclusion complexes are shown in Table 1.

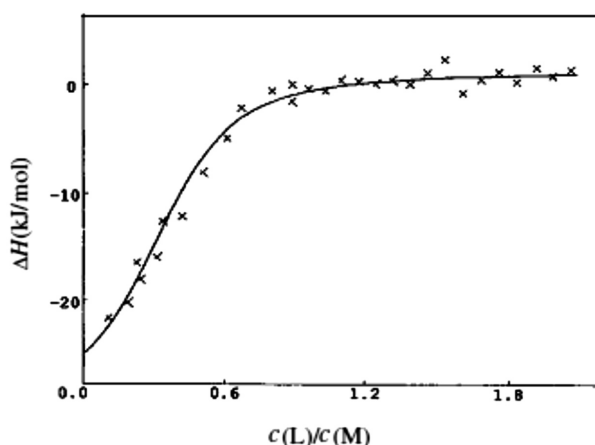


Figure 1 Enthalpy as a function of the concentration ratio $c(L)/c(M)$, where $M = (C_{12}N)_2Cl_2$ and $L = \beta$ -CD for the reactions $M + 2L \rightleftharpoons ML_2$, x, this work and —, calculation.

3.2. Thermodynamic Parameters of the Complexes of β -CD and Gemini Surfactants

The optimum simulated results for the stoichiometry of all (β -CD-(C_{*n*}N)₂Cl₂) association systems are shown in Table 1. Values of β_2 indicate that 2:1 complexes are stable for the (β -CD+(C₁₂N)₂Cl₂/(C₁₄N)₂Cl₂) systems. The complexes can be represented by the general molecular formula, (C_{*n*}N)₂²⁺·(CD)_{*m*}. Since the values of m are all one, it might be believed that the two hydrocarbon tails of a gemini surfactant ion separately interact with the CD molecules, *i.e.* for the β -CD+(C₁₂N)₂Cl₂/(C₁₄N)₂Cl₂ systems, the hydrocarbon chain is sufficiently long that it can coil in the hole of the β -CD molecules as shown in Fig. 3.

The values of the formation enthalpies (ΔH°) of the binding complexes listed in Table 1 are all negative. The values are decided by classical hydrophobic interaction, the expulsion of high enthalpy water molecules from the molecular cavity of the cyclodextrin to the bulk solution, and the destruction of the iceberg structure surrounding the hydrophobic chains of the surfactant ion. We mainly notice the change of ΔH° caused by the increase of the number of -CH₂- groups in hydrocarbon chains of the gemini surfactant. The apparent enthalpic changes per-mol of β -CD ($\Delta H^\circ(\text{CD})$) are -21.6 and -43.1 kJ mol⁻¹, corresponding to the formation of the most stable complexes (C₁₂N)₂Cl₂ (β -CD)₂ and (C₁₄N)₂Cl₂ (β -CD)₂, respectively. Hence, the longer hydrocarbon chain of the surfactant makes the complex more enthalpically stable.

From the thermodynamic formulae, $\Delta G^\circ = -RT \ln \beta_i$ and $T\Delta S^\circ = \Delta H^\circ - \Delta G^\circ$, the standard Gibbs energy (ΔG°) and entropy ($T\Delta S$) changes for the formation of the complexes mentioned above can be derived. The negative Gibbs energy changes (ΔG°) indicate that formation of the (host + guest) inclusion complexes in aqueous solution is a spontaneous process. The influence of the hydrophobic chain on the thermal effect has been discussed above. In order to elucidate the entropy effect one must consider the microstructures involved. The values of $T\Delta S$ for the formation of the β -CD+(C_{*n*}N)₂Cl₂ complexes are negative and decrease from -6.7 to -50.3 kJ mol⁻¹ as the hydrophobic chain

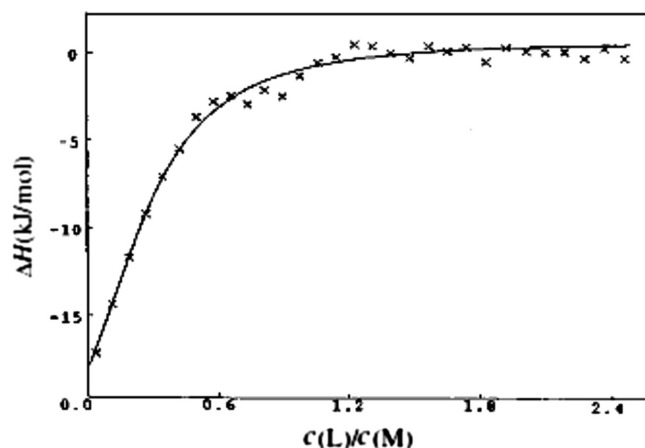


Figure 2 Enthalpy as a function of the concentration ratio $c(L)/c(M)$, where $M = (C_{14}N)_2Cl_2$ and $L = \beta$ -CD for the reactions $M + 2L \rightleftharpoons ML_2$, x, this work and —, calculation.

Table 1 Thermodynamic parameters of complexation for the host (β -cyclodextrin) with the guest ((C_{*n*}N)₂Cl₂) in aqueous solution at 293.15 K.

Surfactants as M	Reaction models	$\Delta H^\circ/\text{kJ mol}^{-1}$	$\beta_2/\text{mol}^{-2} \text{dm}^6$	$\Delta G^\circ/\text{kJ mol}^{-1}$	$T\Delta S^\circ/\text{kJ mol}^{-1}$
(C ₁₂ N) ₂ Cl ₂	$M + 2L \rightleftharpoons ML_2$	-43.10	3.10×10^6	-36.40	-6.70
(C ₁₄ N) ₂ Cl ₂	$M + 2L \rightleftharpoons ML_2$	-86.10	2.40×10^6	-35.80	-50.30

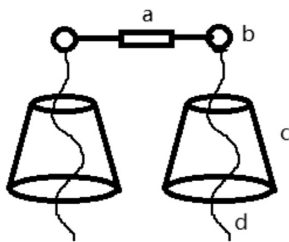


Figure 3 A schematic representation of the inclusion complex (a, spacer group; b, ion head group; c, β -CD; d, hydrophobic chain).

becomes longer. Firstly, the hydrophobic chain of the surfactant can drive water molecules from the hole of the β -CD molecule to the bulk solvent as the inclusion complex is formed. The longer the hydrophobic chain, the more water molecules are driven out from the hole of the CD molecule. This transport of water molecules is an entropy-increasing process. Secondly, collapse of the iceberg structure formed by water molecules around the hydrophobic chain also causes the entropy to increase³¹ when the chain combines with the CD molecule. If merely these two entropy-increasing factors were considered, formation of all the complexes would be accompanied by an increase in entropy. However, there is a factor causing entropy to decrease. That is the hydrophobic interaction between the CD and the hydrophobic tail of the surfactant molecule, which confines the freedom of motion of the host and guest molecules, and so causes entropy to decrease. It is the direct (host + guest) interaction that makes the entropy changes of formation of the (β -CD-(C_nN)₂C1₂) complexes negative.

3.3 Chemical Shifts

The chemical shift variations for protons of the β -CD molecule (Fig. 4) are shown in Fig. 5. The upfield shifts for H3 and H5, induced by microenvironments of high electron density, are a clear sign of the inclusion. With the increase of molar ratio of (C_nN)₂C1₂ to the CD, the chemical shifts of the H3/H5 protons of CD move to a higher field. When the molar ratio of (C_nN)₂C1₂ to CD exceeds 1:2, the change of $\Delta\delta$ is not obvious. It shows that the complexes, in which a molar ratio of CD to (C_nN)₂C1₂



Figure 4 A schematic representation of the H3 and H5 in the β -CD molecule.

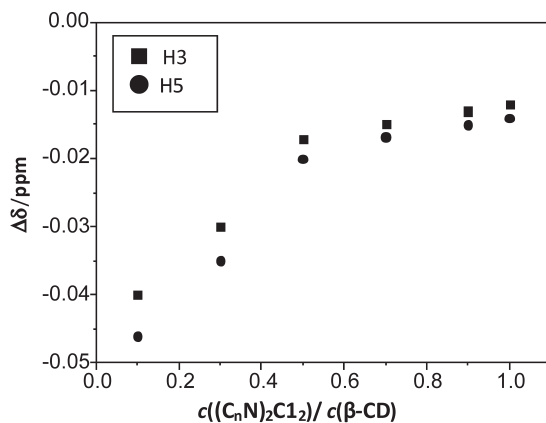


Figure 5 The chemical shift variations for protons in the CD molecule versus the molar ratio of (C_nN)₂C1₂ to β -CD.

exceeds 2:1, cannot form. This result confirms the results derived from the thermodynamic analysis.

4. Conclusions

Isothermal titration calorimetry was applied to determine the binding constants, stoichiometry, and standard enthalpy, entropy, and Gibbs energy changes for complex formation between gemini surfactants, (C_nN)₂C1₂ (where $n = 12, 14$), and β -cyclodextrin (β -CD) in aqueous solution at the temperature of 293.15 K. The systems studied form 2:1 mole ratio host-guest complexes. (Enthalpy + entropy) compensation becomes more evident with an increase of the length of the twin hydrophobic chains in the surfactant molecule. NMR chemical shift data have been determined as evidence of formation for the inclusion complexes and to estimate their stoichiometry.

Acknowledgements

The author thanks the Foundation of LYDX2013BS019 and LYDX2013BS018 for financially supporting this work.

References

- W. Eli, W.H. Chen and Q.J. Xue, *J. Incl. Phenom. Macro*, 2000, **38**, 37–34.
- P. Yan, J.N. Tang and J.X. Xiao, *J. Disper. Sci. Technol.*, 2007, **28**, 617–621.
- Z.W. Gao and X.P. Zhao, *Polymer*, 2003, **44**, 4519–4526.
- H. Xing, S.S. Lin and J.X. Xiao, *J. Chem. Eng. Data*, 2010, **55**, 1940–1944.
- H. Xing, H.T. Zhou, H.Q. Yu, Z.M. Gou and J.X. Xiao, *J. Chem. Eng. Data*, 2011, **56**, 1423–1432.
- D. Wang, P.F. Long, R.H. Dong and J.C. Hao, *Langmuir*, 2012, **28**, 14155–14163.
- J.F. Shi and X.H. Shen, *J. Phys. Chem. B*, 2014, **118**, 1685–1695.
- T. Okubo, H. Kitano and N. Ise, *J. Phys. Chem.*, 1976, **80**, 2661–2664.
- R. Palepu, J.E. Richardson and V.C. Reinsborough, *Langmuir*, 1989, **5**, 218–221.
- W.C. Cromwell, K. Bystrom and M.R. Eftink, *J. Phys. Chem.*, 1985, **89**, 326–332.
- V.K. Smith and T.T. Ndou, *Applied Spectrosc.*, 1992, **46**, 659–668.
- U.R. Dharmawardana, S.D. Christian, E.E. Tucker, R.W. Taylor and J.F. Scamehorn: *Langmuir*, 1993, **9**, 2258–2263.
- M. Tuncay and S.D. Christian, *J. Colloid Interface Sci.*, 1994, **167**, 181–185.
- R.H. Lu, J.C. Hao, H.Q. Wang and L.H. Tong, *J. Inclusion Phenom. Mol. Recognit. Chem.*, 1997, **192**, 213–221.
- J. Czapkiewicz and B. Tutaj, *J. Inclusion Phenom. Mol. Recognit. Chem.*, 1993, **16**, 377–382.
- N. Funasaki, S. Ishikawa and S. Neya, *J. Phys. Chem. B*, 2003, **107**, 10094–10099.
- P.M. Sehgal, R. Wimmer, K.L. Larsen and D. Otzen, *Colloid Polym Sci.*, 2006, **284**, 916–926.
- Y. Bai, G.Y. Xu, X. Xin and H.Y. Sun, *Colloid Polym. Sci.*, 2008, **286**, 1475–1484.
- D. Geetha, T. Thilagavathi, P.S. Ramesh and C. Rakkappan, *J. Engn. Appl. Sci.*, 2009, **4**, 60–65.
- S.K. Hait and S.P. Moulik, *Curr. Sci.*, 2002, **82**, 1101–1111.
- O. Rist and A. Rike, *Molecules*, 2001, **6**, 979–987.
- T.S. Kim, T. Hirao and I. Ikeda, *J. Am. Oil Chem. Soc.*, 1996, **73**, 67–71.
- C.G. Ricci, M.I. Cabreera, J.A. Luna and R.J. Grau, *Synlett.*, 2002, **11**, 1811–1814.
- A.M. Jacobson and Z. Casassa, *J. Coll. Interf. Sci.*, 1991, **142**, 489–496.
- M.S. Wikinson and R.L. Kidwell, *J. Coll. Interf. Sci.*, 1971, **35**, 114–118.
- G. Bai, Y. Wang and H. Yan, *J. Phys. Chem. B*, 2002, **106**, 2153–2159.
- F.T. Gucker Jr. and R.W. Planck, *J. Am. Chem. Soc.*, 1939, **61**, 459–470.
- Li.D. Song and M.J. Rosen, *Langmuir*, 1996, **12**, 1149–1153.
- I. Gómez-Orellana and D. Hallén, *Thermochim. Acta*, 1993, **221**, 183–193.
- D.Z. Sun, S.B. Wang, X.L. Wei and B.L. Yin, *J. Chem. Thermodyn.*, 2005, **37**, 431–436.
- M.V. Rekharsky and Y. Inoue, *Chem. Rev.*, 1998, **98**, 1875–1917.